91. The element with atomic number Z=118 will be :–
(1) Noble gas (2) Transition metal (3) Alkali metal (4) Alkaline earth metal

92. In the general electronic configuration -
$(n-2)f^{1-14}\ (n-1)d^{0-1}\ ns^2$, if value of n = 7 the configuration will be of -
(1) Lanthenides (2) Actinides (3) Transition elements (4) Boble gas

93. Diagonal relationship is shown by
(1) Elements of second period
(2) Elements of third period
(3) Both (1) and (2)
(4) Elements of First period

94. In the ions P^{3-} , S^{2-} and Cl^- the increasing order of size is:-
(1) $Cl^- < S^{2-} < P^{3-}$ (2) $P^{3-} < S^{2-} < Cl^-$ (3) $S^{2-} < Cl^- < P^{3-}$ (4) $S^{2-} < P^{3-} < Cl^-$

95. Atomic radii of Fluorine and Neon in Angstrom units are given by :-
(1) 0.72, 1.60 (2) 1.60, 1.60 (3) 0.72, 0.72 (4) 1.60, 0.72

96. Which of these have no unit?
(1) Electronegativity (2) Electron affinity
(3) Ionisation energy (4) Atomic radius

97. A sudden large jump between the values of 2^{nd} and 3^{rd} IP of an element would be associated with the electronic configuration :-
(1) $1s^2, 2s^2\ 2p^6, 3s^1$
(2) $1s^2, 2s^2\ 2p^6, 3s^2\ 3p^5$
(3) $1s^2, 2s^2\ 2p^6, 3s^2\ 3p^2$
(4) $1s^2, 2s^2\ 2p^6\ 3s^2$

98. Compared to the first ionisation potential, the value of second ionisation potential of an element is :-
(1) Negligible (2) Smaller (3) Greater (4) Double

99. In the process $Cl_{(g)} + e^- \xrightarrow{\Delta H} Cl^-(g)$, ΔH is
(1) Positive (2) Negative
(3) Zero (4) None

100. Electronegativity values for elements are useful in predicting :-
(1) Bond energy of a molecule
(2) Polarity of a molecule
(3) Nature of an oxide
(4) All