

61. Electronic configuration of chalcogens in their outermost orbit is
 (1) $s^2 p^3$ (2) $s^2 p^4$ (3) $s^2 p^5$ (4) $s^2 p^6$
62. The electronic configuration of the element which is just above the element with atomic number 43 in the same groups is...
 (1) $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^5 4s^2$
 (2) $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^5, 4s^2 4p^6$
 (3) $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^6 4s^2$
 (4) $1s^2, 2s^2 2p^5, 3s^2 3p^6 3d^7 4s^2$
63. Match the Column I with Column II and select the correct answer by given codes.
Column I (Atomic number)
 A. 52
 B. 56
 C. 57
 D. 60
Column II (Position in the Periodic Table)
 1. s-block
 2. p-block
 3. d-block
 4. f-block
- | | A | B | C | D |
|-----|---|---|---|---|
| (1) | 2 | 1 | 3 | 4 |
| (2) | 2 | 1 | 4 | 3 |
| (3) | 1 | 2 | 3 | 4 |
| (4) | 1 | 2 | 4 | 3 |
64. On moving from left to right across a period in the table the metallic character
 (1) Increases (2) Decreases
 (3) Remains constant (4) First increases and then decreases
65. The given reaction, $X(g) + e^- \rightarrow X^-(g)$ is an example of
 (1) electron gain enthalpy (2) electron loss enthalpy
 (3) electron gain energy (4) Both (1) and (3)
66. The pair of neutral oxides
 (1) CO, CO₂, NO₂ (2) NO, NO₂, CO₂ (3) Both (1) and (2) (4) CO, NO, NO₂
67. Electronegativity is inversely related to which the following property.
 (1) non-metallic properties (2) metalloid properties
 (3) ionic properties (4) metallic properties
68. Among the elements Ca, Mg, P and Cl, the order of decreasing atomic radii is
 (1) Ca > Mg > P > Cl (2) P > Cl > Ca > Mg
 (3) Mg > Ca > Cl > P (4) Cl > P > Mg > Ca
69. Consider the isoelectronic species, Na^+, Mg^{2+}, F^- And O^{2-} . The correct order of their increasing ionic radii is
 (1) $F^- < O^{2-} < Mg^{2+} < Na^+$ (2) $Mg^{2+} < Na^+ < F^- < O^{2-}$
 (3) $O^{2-} < F^- < Na^+ < Mg^{2+}$ (4) $O^{2-} < F^- < Mg^{2+} < Na^+$
70. The normal oxide formed by the element present on extreme left in the Periodic Table are ...A..., present. Here, A refers to
 (1) most acidic (2) most basic (3) Amphoteric (4) Both (1) and (2)