61 (2) Chalcogens are oxygen family elements i.e., group 16 elements.

62 (1) Atomic number of 43 is known as the name of technetium and the element, which is just above the technetium atom is manganese (Mn) atomic number = 25 Electronic configuration of Mn = $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^5 4s^2$

63 (1) $A.\ 52 \rightarrow ns^2 np^{1-6}(p-block)$

$B. 56 \rightarrow ns^{1-2}(s-block)$

$C. 57 \rightarrow (n-1)d^{1-10} ns^{1-2}(d-block)$

$D. 60 \rightarrow (n-2)f^{1-14}(n-1)d^{0-1}ns^2\ (f-block)$

64 (2) Metallic Character decreases as the electropositive character decreases

65 (4) Electron gain enthalpy provides a measure of the ease with which an atom adds an electron to from anion as repesented by.

X(g) + e^-(g) $\rightarrow X^-$

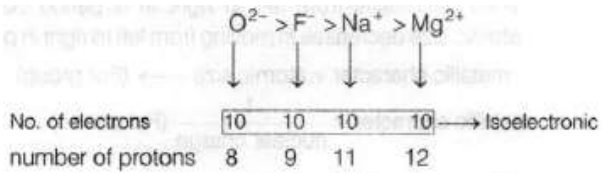
Electron gain enthalpy = electron gain energy

66 (4)

67 (4) Electrogtivity (EN) $\alpha \frac{1}{Metallic\ properties}$ En decrases down a group is accompanied by a decreases in non metllic properties of elements.

68 (1) Atomic radii increase down the group while decrease from left to right in a period. So, Ca> Mg> P> CI is order of decreeasing atomic radii

69 (2) In case of isoelectronioc species ionic radii decrease as number of protonos increases.



70 (2) The normal oxide formed by the elements present on the extreme left of Periodic. Table is the most basic e,g: Na_2O whereas that formed by the elements present on extreme right to the Periodic Tableisthemos acidic.

e.g: Cl_2O_1