

CHEMISTRY**Periodic Properties****Electron Affinity**

41. To which of the following atom, addition of electron is most difficult?
- (1) Radon (2) Nitrogen
(3) Oxygen (4) Radium
42. Which of the following processes involves absorption of energy?
- (1) $S(g) + e^- \rightarrow S^-(g)$
(2) $S^- + e^- \rightarrow S^{2-}(g)$
(3) $Cl(g) + e^- \rightarrow Cl^-(g)$
(4) none of these
33. Among the following configurations, the element which has the highest electron affinity is
- (1) $[Ne]3s^13p^2$
(2) $[Ne]3s^23p^5$
(3) $[Ne]3s^23p^4$
(4) $[Ne]3s^23p^63d^54s^1$
44. The increasing order of electron affinity of the electronic configurations of element is
- (I) $1s^22s^22p^63s^23p^5$
(II) $1s^22s^22p^3$
(III) $1s^22s^22p^5$
(IV) $1s^22s^22p^63s^23p^3$
- (1) II < IV < III < I
(2) I < II < III < IV
(3) I < II < III < IV
(4) IV < III < II < I
35. Second electron gain enthalpy
- (1) is always negative
(2) is always positive
(3) can be positive or negative
(4) is always zero
46. The element having very high ionisation enthalpy but zero electron affinity is
- (1) H (2) F
(3) He (4) Be
47. The process requiring absorption of energy is
- (1) $N \rightarrow N^-$ (2) $F \rightarrow F^-$
(3) $Cl \rightarrow Cl^-$ (4) $H \rightarrow H^-$
38. Which of the following is the correct order for electron gain enthalpy?
- (1) $P < O < S < F$
(2) $O < P < S < F$
(3) $F < S < P < O$
(4) $P < S < O < F$
39. The second electron gain enthalpies (in kJ/mol) of oxygen and sulphur respectively are
- (1) -780, +590 (2) -590, +780
(3) +590, +780 (4) +780, +590
40. Which of the following is correct order of electron affinity?
- (1) $Be > Mg > Ca > Sr > Ba$
(2) $N > P > As > Sb > Bi$
(3) $Ne > Ar > Kr > Xe > Rn$
(4) $Cl > F > Br > I$