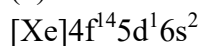


CHEMISTRY
Periodic Properties**Periodic Table**

1. (1)
2. (3)
3. (4)
4. (4)
5. (1)
6. (4)
7. (2)

Tantalum ($Z = 73$) is a transition element.

8. (3)



Atomic number = 71

Period number = 6th

Group number = 3rd

9. (3)

10. (3)

P is trivalent non-metal Q is divalent metal hence formula of compound is



CHEMISTRY
Periodic Properties**Atomic/Ionic Radius**

11. (4)

12. (2)

13. (2)

14. (2)

15. (2)

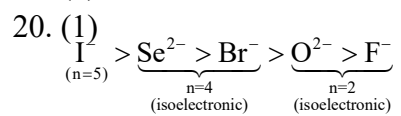
16. (2)

17. (3)

18. (2)

19. (2)

20. (1)

**PARISHRAMA
NEET ACADEMY**

NEET 220 Medical Seats out of 240

CHEMISTRY

Periodic Properties

Ionisation Energy

21. (4)
22. (3)
23. (2)
24. (3)
25. (2)
26. (2)
27. (4)
28. (2)

High jump between IE_4 and IE_5

Therefore, number of valence = 4

General configuration ns^2np^2

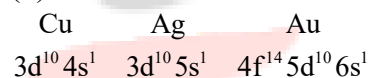
Therefore element = Si

29. (4)

For lower value of n , Z_{eff} is high therefore

IE required will also be high.

30. (1)



In Au, poor screening is provided by 4f subshell electron.

$\therefore Z_{\text{eff}}$ increases and ionisation energy also increases.

$\therefore Au > Cu > Ag$ is correct order.

CHEMISTRY
Periodic Properties**Electron Affinity**

31. (1)

32. (2)

33. (2)

34. (1)

35. (2)

36. (3)

37. (1)

Nitrogen has stable $2p^3$ configuration and also due to high electron charge density at outermost orbital it requires energy to add one extra electron in its outer most shell i.e., its first electron gain enthalpy is positive.

38. (1)

$P < O < S < F$ (correct order of electron affinity)

39. (4)

Successive electron gain enthalpies is always positive.

O^- has higher positive value of electron gain enthalpy than S^- because more energy required to overcome the electronic repulsion between small sized O^- ion and the incoming electron.

40. (4)

Correct order of E.A. is

$(Be < Mg < Ca < Sr < Ba)$,

$(N < P < As < Sb < Bi)$,

$(Ne < Ar < Kr < Xe < Rn)$,

$(Cl > F > Br > I)$.

CHEMISTRY

Periodic Properties

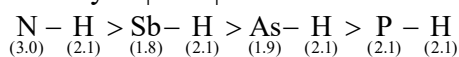
Electronegativity

41. (2)

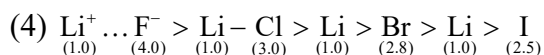
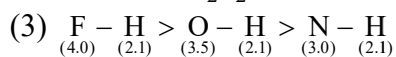
42. (3)

$$\chi_p = \frac{\chi_m}{2.8}; \text{ so } \chi_m > \chi_p$$

43. (2)

Polarity $\propto |\Delta EN|$ 

(2) EN of carbon atom in $\text{CH}_2\text{F}_2 > \text{CH}_4$
due to more positive charge on carbon
atom in CH_2F_2 .



PARISHRAMA NEET ACADEMY