

CHEMISTRY

41. (1)

42. (1)

$$1 \text{ mol of O}_3 = 22.4 \text{ L}$$

$$0.2 \text{ mol of O}_3 = 22.4 \times 0.2$$

$$1 \text{ mol O}_3 = 48 \text{ g}; 0.2 \text{ mol O}_3 = 0.2 \times 48 = 9.6 \text{ g}$$

43. (4)

$$98.892 \times 12$$

$$+ 1.108 \times 13$$

$$\text{Average atomic weight} = \frac{+14 \times 2 \times 10^{-10}}{100} = 12.0114$$

44. (2)

$$\text{Mass} = 2.5 \times 2.15 = 5.375 \text{ g}$$

$$\text{In correct S.F, mass} = 5.4 \text{ g}$$

45. (4)

46. (1)

47. (4)

48. (1)

49. (4)

50. (4)

$$e^- = \frac{4.2}{14} \times 10 \times N_A$$

**PARISHRAMA
NEET ACADEMY**