

CHEMISTRY

31. Mass of one CO molecule in gram is
(1) 4.65×10^{-23} (2) 1.66×10^{-24}
(3) 3×10^{-24} (4) 6.22×10^{-23}
32. Given that one mol of N_2 at NTP occupies 22.4 L the density of N_2 is
(1) 1.25 g L^{-1} (2) 0.80 g L^{-1}
(3) 2.5 g L^{-1} (4) 1.60 g L^{-1}
33. A gaseous mixture of CH_4 and O_2 contains equal masses of both. The mole ratio in the mixture is
(1) $\frac{1}{2}$ (2) $\frac{1}{3}$
(3) $\frac{2}{1}$ (4) $\frac{1}{4}$
34. 0.0014 can be written in scientific notation as
(1) 0.14×10^{-2} (2) 1.4×10^{-3}
(3) 14×10^{-3} (4) 140×10^{-3}
35. Law of constant composition does not hold good in
(1) Ionic hydrides
(2) Covalent hydrides
(3) Metallic hydrides
(4) All of these
36. Majority of H_2 gas is produced from _____
(1) petrochemicals
(2) electrolysis of aqueous solutions
(3) zinc dust
(4) coal
37. $NaAlSiO_4$ is known as
(1) calgon (2) zeolite
(3) lime water (4) resin
38. The normality of 10 vol H_2O_2 is nearly
(1) 2.1 (2) 3.4
(3) 1.7 (4) 5.1
39. Maximum number of hydrogen bonds per molecule in water are
(1) 1 (2) 2
(3) 3 (4) 4
40. 12.1706 can be rounded up in three significant figures as _____
(1) 12.2 (2) 12.8
(3) 17.21 (4) 12.72