

CHEMISTRY

31. (1)

1 mol of CO = 28 g = Avogadro number of molecules

$$\text{One molecule CO} = \frac{28}{6 \times 10^{23}} \text{ g} = 4.65 \times 10^{-23} \text{ g}$$

32. (1)

$$\therefore d = \frac{m}{V} \quad 1 \text{ mol of N}_2 = 28 \text{ g, } V = 22.4 \text{ L}$$

$$\therefore \text{density} = \frac{28}{22.4} \text{ g L}^{-1} = 1.25 \text{ g L}^{-1}$$

33. (3)

$$\frac{n_{\text{CH}_4}}{n_{\text{O}_2}} = \frac{1}{1} = \frac{2}{1}$$

34. (2)

35. (3)

36. (1)

37. (2)

38. (3)

10 vol H₂O₂, volume strength = 3%

$$\text{Normality} = \frac{\text{Gram equivalent of solute}}{\text{Volume of solute}}$$

$$= \frac{17}{10} = 1.7$$

39. (4)

40. (1)



PARISHRAMA
NEET ACADEMY