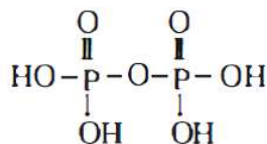
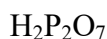


CHEMISTRY

p-Block Elements (from 13 to 18 Group Elements)

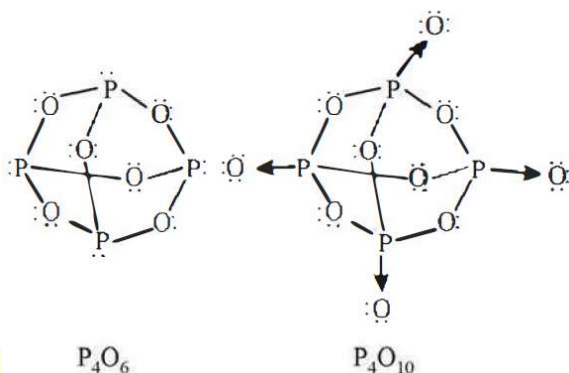
11. (2)



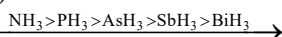
Tetrabasic

4-OH groups are present

12. (1)



13. (2)



Stability decreases down the group because bond energy decreases down the group.

14. (4)

	I	II	III	IV	V
Element	N	P	As	Sb	Bi
Atomic number	7	15	33	51	83

15. (1)

Hypophosphorous acid (H_3PO_2) is a monobasic acid which act as reducing agent. In this molecule, two P–H bonds are responsible for its reducing character and one –OH group is responsible for its monobasic acid character.

16. (2)

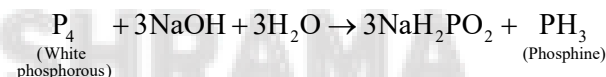
The structure of NO is $\ddot{\text{N}} = \dot{\text{O}}$. It is an odd electron molecule and is paramagnetic. It contains a three-electron bond.

17. (4)

He has 2 electrons in its outermost shell.

18. (2)

The reaction between NaOH and white phosphorous (P_4) can be represented as follows



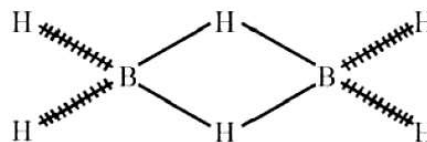
In this reaction phosphorous is oxidised as well as reduced so it is a disproportionation reaction.

19. (1)

Boric acid is a weak monobasic acid as it is not able to release H^+ ions on its own. It receives OH^- ions from water to complete its octet and in turn releases H^+ ions.

20. (4)

B_2H_6 has two types of B–H bonds



B 119pm H (Terminal bond)

B 134pm H (Bridge bond)