

CHEMISTRY
s-Block Elements**General Characteristics of Group 1 Elements**

1. (4)
The electrons are excited to higher energy level and on return to original state they give out visible complex formation.
2. (1)
Because of smallest size, metallic bonding is the strongest in lithium.
3. (4)
The atomic numbers of alkali metals are: 3, 11, 19, 37, 55 and 87.
4. (4)
As the size increases, the metallic bonding decreases.
5. (2)
Alkali metals are strong reducing agents and have less value for I.P.
6. (3)
A characteristic feature of Na-K alloy.
7. (2)
Lithium has the lowest density (0.534 g cm^{-3})
8. (4)
Potassium salts (i.e. KCl) impart pink violet colour the flame.
9. (4)
One electron in the outer most shell.
10. (1)
In any period, alkali metal (except noble gas) has the lowest density. Therefore alkali metals have large atomic volumes.

 **PARISHRAMA
NEET ACADEMY**